

Chapter 6-7 Jeopardy Game

| | Value | Question | Answer |
|-----------------|-------|--|---------------------------------------|
| Classification | 200 | This is the name of the group 1 elements. | Alkali metals |
| | 400 | Most elements are what type of material? | Metals |
| | 600 | Elements from the s- and p-block are called what? | Representative |
| | 800 | These elements have properties of metals and nonmetals. | Metalloids |
| | 1000 | What are the names of the two rows that make up the inner transition metals? | Lanthanoids & Actinoids |
| Periodic Trends | 200 | This is the smallest element on the periodic table. | He |
| | 400 | This element has the smallest ionization energy. | Fr |
| | 600 | Of Si, S, or P, the one with the highest electronegativity | S |
| | 800 | Which is larger, As^{3-} or As? | As^{3-} |
| | 1000 | Which atomic radius is largest: S, Cl, or Br? | Not enough info (trend reversed) |
| Properties | 200 | These elements are dull, brittle, and non-conductive | Non-metals |
| | 400 | These are generally hard, strong metals | Transition Metals |
| | 600 | How many electrons are needed to fill the outermost energy level? | 8 |
| | 800 | This is the name for the energy needed to remove an electron from an atom. | Ionization Energy |
| | 1000 | Why were the rare earth metals difficult to separate and discover? | Similar P/C Properties |
| Families | 200 | This is the name of the group 18 elements | Noble Gases |
| | 400 | All elements in a family have the same what? | # of Valence Electrons |
| | 600 | These nonmetal elements react with metals to form salts. | Halogens |
| | 800 | Group 13 elements form ions with what charge? | 3+ |
| | 1000 | These elements are harder and slightly less reactive than group IA. | Alkaline earth metals |
| Regions | 200 | This block encompasses the alkali metals and alkaline earth metals | s-block |
| | 400 | This is the region above and to the right of the staircase | non-metals |
| | 600 | What elements exist along the staircase on the right side of the table? | Metalloids |
| | 800 | What is another name for the f-block elements? | Inner Transition or rare earth metals |
| | 1000 | All gaseous elements belong to what category? | non-metals. |
| Potpourri | 200 | How does an atom become positively charged? | Loses an electron. |
| | 400 | What is the charge on a nickel (Ni) ion with electron configuration $[Ar]3d^7$? {Hint: what is the electron configuration of a neutral Ni atom?} | 3+ |
| | 600 | Diamond is one what of carbon? | Allotrope |
| | 800 | All Actinoid (5f) elements share what dangerous property? [Hint: what is true of all elements with atomic number $Z > 83$?] | Radioactivity |
| | 1000 | This energy level of d-orbitals is being filled in the 6th row. | 5 |

Final Jeopardy Category: Transition metals

These are responsible for the magnetic properties of the transition metals and the color of their compounds.

Answer: Unpaired d-electrons