

WKS

**Formula & Nomenclature Practice**  
**Binary & Ternary Ionic Compounds**

NAME \_\_\_\_\_

Period \_\_\_\_\_ Date \_\_\_\_\_

<u>Name to Formula</u> Put charges on ions first, then make neutral. Omit charges (but keep subscripts) on final answer.	<u>Formula to Name</u> Name the cation and the anion.
1. calcium fluoride	13. MgBr <sub>2</sub>
2. aluminum oxide	14. CsF
3. magnesium nitride	15. Al <sub>2</sub> S <sub>3</sub>
4. sodium phosphide	16. NH <sub>4</sub> Cl
5. magnesium hydroxide	17. LiCN
6. sodium hydrogen sulfate	18. Rb <sub>3</sub> P
7. strontium oxide	19. Mg(HCO <sub>3</sub> ) <sub>2</sub>
8. aluminum arsenide	20. NaF
9. calcium hypochlorite	21. Na <sub>3</sub> N
10. strontium carbonate	22. K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub>
11. ammonium sulfite	23. NaBrO <sub>2</sub>
12. barium nitrate	24. KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>

25. How do you determine the correct subscripts in a chemical formula?

26. What is the difference between a monatomic ion and a polyatomic ion? Give an example of each.

27. What is an oxyanion and what do the different endings indicate?