

**WKS**  
**Formula & Nomenclature Practice**  
**Transition Metal & Mixed Ionic Compounds**

NAME \_\_\_\_\_  
Period \_\_\_\_\_ Date \_\_\_\_\_

<u>Name to Formula</u> (Put charges on ions first!!! Then make neutral.)	<u>Formula to Name</u> (Roman numerals when needed!)
1. tin(IV) chromate	16. PbO <sub>2</sub>
2. zinc carbonate	17. FeSO <sub>4</sub>
3. mercury (II) chloride	18. Ni(NO <sub>2</sub> ) <sub>3</sub>
4. cobalt(II) sulfide	19. Mg <sub>3</sub> P <sub>2</sub>
5. iron (III) oxide	20. PdCO <sub>3</sub>
6. silver cyanide	21. CuCl <sub>2</sub>
7. calcium oxalate (don't reduce anion!)	22. CdSe
8. vanadium(V) hydroxide	23. Mo <sub>2</sub> S <sub>3</sub>
9. barium sulfite	24. NbI <sub>5</sub>
10. manganese (IV) carbonate	25. Li <sub>3</sub> PO <sub>4</sub>
11. copper (II) nitrate	26. KCl
12. strontium iodide	27. RuBr <sub>3</sub>
13. ammonium sulfide	28. Zn(OH) <sub>2</sub>
14. chromium(VI) oxide	29. BeSO <sub>4</sub>
15. iridium(IV) phosphate	30. AgNO <sub>3</sub>

31. What does the Roman numeral in the name of a cation indicate? How can it be determined from the formula?

32. When is it necessary to use a Roman numeral in the name of a cation? What cations do NOT need a Roman numeral?