WKS: Additional
Lewis Structures-VSEPR

NAME _____
Period _____ **Date** _____

Determine the number of valence electrons and draw the Lewis structures for the following molecules or polyatomic ions. Determine the Electron & Molecular Geometries using VSEPR, then draw the 3D structure. Lewis structures must contain ALL valence electrons, but the 3D drawing does not need the lone pairs on single-bonded terminal atoms.

Formula & # val. e^-	Lewis Structure (Show ALL electrons)	Electron & Molecular Geometries	3D Drawing (Omit lone pairs on single-bonded terminal atoms)	Bond Angle
1. H_2S				
2. CHFO (C is central)				
3. NO_3^-				
4. NBr_3				
5. CH_3I				
6. SiO_3^{2-}				

[over]

Formula & # val. e^-	Lewis Structure (Show ALL electrons)	Electron & Molecular Geometries	3D Drawing (Omit lone pairs on single-bonded terminal atoms)	Bond Angle
7. N_2O (Arranged N-N-O)				
8. SiO				
9. BeF_2*				
10. BCl_3*				
11. SO_3^{2-}				
12. ClNO (Cl-N-O)				
13. SCO (C central)				

*B & Be are exceptions to the Octet Rule: they can be "satisfied" with fewer electrons.