

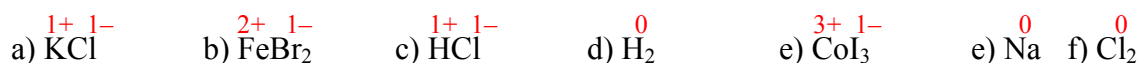
1) When a substance is oxidized, its charge (**increases, decreases**). Thus, when a substance is oxidized it (**gains, loses**) electrons.

2) When a substance is reduced, its charge (**increases, decreases**). Thus, when a substance is reduced it (**gains, loses**) electrons.

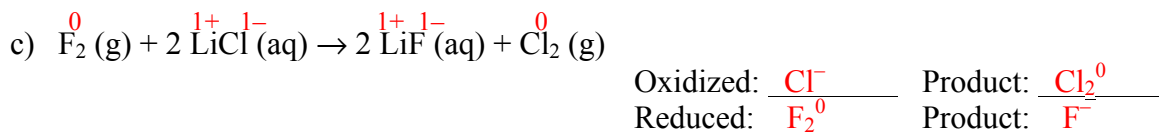
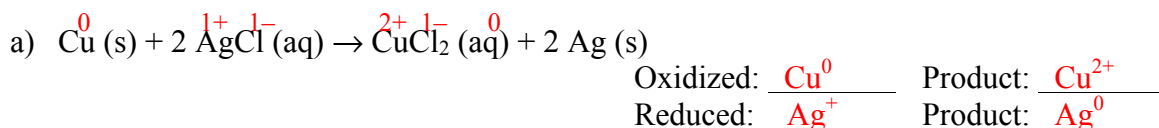
3) Label these changes as either oxidation or reduction:



4) Put in all charges (oxidation numbers) in the following substances. (Remember: The charge on pure elements is zero. Also, when H is in a compound, its charge is +1)



5) Put in all charges (oxidation numbers). Then indicate which substance is being oxidized, which is being reduced, and what their products are. [Remember LEO-GER!]



6) Complete the following reactions. Put in all charges (oxidation numbers). Then indicate which substance is being oxidized, which is being reduced, and what their products are. [Remember LEO-GER!]

