

Chart E: Activity Series

Metals and Hydrogen (Cations)		
Least Reactive ↑ Most Reactive	Li (s) → Li ⁺ (aq) + e ⁻	React with H ⁺ to form H ₂ in H ₂ O or acid
	K (s) → K ⁺ (aq) + e ⁻	
	Ba (s) → Ba ²⁺ (aq) + 2e ⁻	
	Sr (s) → Sr ²⁺ (aq) + 2e ⁻	
	Ca (s) → Ca ²⁺ (aq) + 2e ⁻	
	Na (s) → Na ⁺ (aq) + e ⁻	
	Mg (s) → Mg ²⁺ (aq) + 2e ⁻	
	Al (s) → Al ³⁺ (aq) + 3e ⁻	React with H ⁺ to form H ₂ in acid only
	Mn (s) → Mn ²⁺ (aq) + 2e ⁻	
	Zn (s) → Zn ²⁺ (aq) + 2e ⁻	
	Cr (s) → Cr ²⁺ (aq) + 2e ⁻	
	Fe (s) → Fe ²⁺ (aq) + 2e ⁻	
	Cd (s) → Cd ²⁺ (aq) + 2e ⁻	
	Co (s) → Co ²⁺ (aq) + 2e ⁻	
	Ni (s) → Ni ²⁺ (aq) + 2e ⁻	
	Sn (s) → Sn ²⁺ (aq) + 2e ⁻	Do not react with acid or water to form H ₂
	Pb (s) → Pb ²⁺ (aq) + 2e ⁻	
	H ₂ (g) → 2H ⁺ (aq) + 2e ⁻	
	Sb (s) → Sb ²⁺ (aq) + 2e ⁻	
	Bi (s) → Bi ²⁺ (aq) + 2e ⁻	
Cu (s) → Cu ²⁺ (aq) + 2e ⁻		
Ag (s) → Ag ⁺ (aq) + e ⁻		
Pd (s) → Pd ²⁺ (aq) + 2e ⁻		
Hg (ℓ) → Hg ²⁺ (aq) + 2e ⁻		
Pt (s) → Pt ²⁺ (aq) + 2e ⁻		
Au (s) → Au ³⁺ (aq) + 3e ⁻		