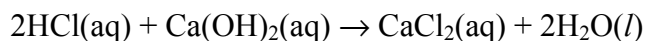


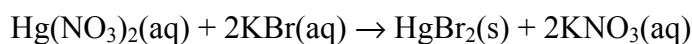
WKS
Solution Stoichiometry I

Name _____
Period _____

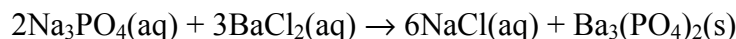
1. How many mL of 0.25 M $\text{Ca}(\text{OH})_2$ are needed to completely react with 10.0 mL of 0.25 M HCl ? The balanced reaction is:



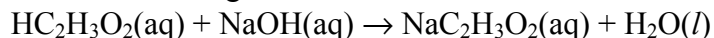
2. How many mL of 0.348 M $\text{Hg}(\text{NO}_3)_2$ are needed to fully react with 15.0 mL of 0.485 M KBr according to the following balanced reaction?



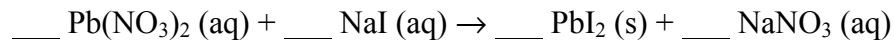
3. What volume of 0.325 M Na_3PO_4 would be needed to fully react with 25.0 mL of 0.480 M BaCl_2 by the following balanced reaction?



4. Calculate the volume of 0.2250 M $\text{HC}_2\text{H}_3\text{O}_2$ (acetic acid) solution needed to neutralize (use up) 25.19 mL of 0.4295 M NaOH in the following balanced reaction.



5. How many mL of 0.120 M $\text{Pb}(\text{NO}_3)_2$ are required to react with 75.0 mL of 0.280 M NaI according to the following equation? You must balance the equation first!



6. How many mL of 0.280 M barium nitrate solution are required to fully react with 25.0 mL of 0.350 M aluminum sulfate solution to form solid barium sulfate and aluminum nitrate solution? Write and balance the reaction first.

Answers: 1) 5.00 mL $\text{Ca}(\text{OH})_2$; 2) 10.5 mL $\text{Hg}(\text{NO}_3)_2$; 3) 0.0246 L Na_3PO_4 ; 4) 48.08 mL $\text{HC}_2\text{H}_3\text{O}_2$; 5) 1.2, 1.2; 87.5 mL $\text{Pb}(\text{NO}_3)_2$; 6) 3, 1.3, 2; 93.8 mL barium nitrate