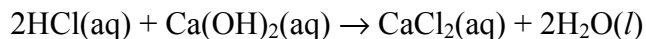


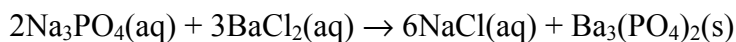
WKS
Solution Stoichiometry I

Name _____
Period _____

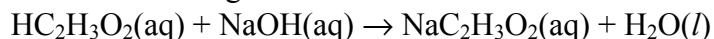
1. How many mL of 0.25 M $\text{Ca}(\text{OH})_2$ are needed to completely react with 10.0 mL of 0.25 M HCl? The balanced reaction is:



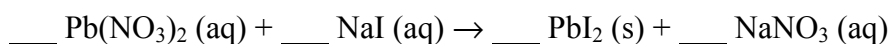
2. What volume of 0.325 M Na_3PO_4 would be needed to fully react with 25.0 mL of 0.480 M BaCl_2 by the following balanced reaction?



3. Calculate the volume of 0.2250 M $\text{HC}_2\text{H}_3\text{O}_2$ (acetic acid) solution needed to neutralize (use up) 25.19 mL of 0.4295 M NaOH in the following balanced reaction.

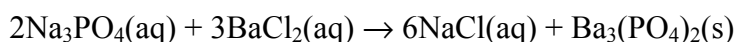


4. How many mL of 0.120 M $\text{Pb}(\text{NO}_3)_2$ are required to react with 75.0 mL of 0.280 M NaI according to the following equation? You must balance the equation first!

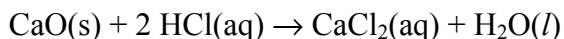


5. How many mL of 0.280 M barium nitrate solution are required to fully react with 25.0 mL of 0.350 M aluminum sulfate solution to form solid barium sulfate and aluminum nitrate solution? Write and balance the reaction first.

6. What volume of 0.325 M Na_3PO_4 would be needed to precipitate 25.00 g $\text{Ba}_3(\text{PO}_4)_2$ with excess BaCl_2 by the following balanced reaction?



7. How many grams of CaO are required for complete reaction with the HCl in 275 mL of a 0.523 M HCl solution? The balanced equation for the reaction is:



8. How many mL of 0.750 M cobalt (III) nitrate, when added to a sodium sulfate solution, are needed to precipitate 8.07 g of cobalt (III) sulfate? Balance the equation first.

