

WKS
Strengths of Acids and Bases

Name _____
Period _____ Date _____

- 1) What is the amount of ionization in strong acids & bases?
- 2) What are the six strong acids?
- 3) What is $[H^+]$ (or $[H_3O^+]$) of a solution of 0.25 M HNO_3 ? How do you know?
- 4) What must be true of all other acids? How do we describe the amount of their ionization? What happens to the strength of an acid as the strength of the H–X or O–H bond *increases*?
- 5) How do we compare the relative strengths of weak acids?
- 6) Use **Table 19-2** (pg. 605) to predict which aqueous solution would have the greater electrical conductivity: 0.1 M $HClO$ or 0.1 M HF . Explain.
- 7) What compounds comprise the strong bases?
- 8) What is $[OH^-]$ of a solution that is 0.075 M $Ba(OH)_2$ (watch the stoichiometry!)?
- 9) Aniline ($C_6H_5NH_2$) and ethylamine ($CH_3CH_2NH_2$) are both weak bases. According to **Table 19-4** (pg. 607), which one is more ionized in aqueous solution? Explain your reasoning.