

Dr. Casagrande's Chemistry 1 Final Exam Review Sheet 3

Chapter 17: Reaction Rates

- Relate rates of chemical reactions to collisions between particles.
- Describe a reaction energy diagram and identify the activation energy and the activated complex (transition state).
- Identify the factors we have observed that affect the rates of chemical reactions.
- Explain the affect of a catalyst and inhibitor.
- Lab: Rates of Reactions.

Chapter 18: Chemical Equilibrium

- Explain what a reversible reaction is and write the equation for a reversible reaction.
- Recognize and describe the characteristics of a system at equilibrium.
- Write equilibrium expressions (K_{eq}) for systems that are at equilibrium (remember no solids and liquids).
- Calculate equilibrium constants from concentration data.
- Reading: What's so Equal about Equilibrium?

Chapter 19: Acids and Bases

- Identify the physical and chemical properties of acids and bases.
- Classify solutions as acidic, basic, or neutral based on the relative levels of $[H^+]$ and $[OH^-]$.
- Recall the Arrhenius (A) and Brønsted (B) models for acids and bases: acids provide (donate) H^+ ions (A & B) and bases either provide OH^- ions (A) or accept an H^+ from an acid (B).
- Identify the hydrogen atom in a polar bond that is ionized.
- ~~• Determine whether an acid is mono-, di-, or triprotic.~~
- Identify the acid, base, conjugate acid, and conjugate base in a reaction.
- Relate the strength of an acid or base (strong or weak) to its degree of ionization (fully or partially ionized) and its other properties, such as electrical conductivity.
- Write an acid ionization (dissociation) reaction equation ~~and write the acid ionization constant expression for it.~~
- Compare the strengths of weak acids or bases from the values of their acid or base ionization constants (K_a or K_b).
- Explain how the strength of an acid or base relates to its strength as an electrolyte.
- Given the $[H^+]$ or $[OH^-]$, calculate the other from $K_w = [H^+][OH^-] = 1.0 \times 10^{-14}$.
- Given the $[H^+]$ calculate pH or given the $[OH^-]$ calculate pOH.
- Classify solutions as acidic, neutral, or basic based on their pH.
- Given pH or pOH, determine the other from $pH + pOH = 14$.
- Describe what an acid-base neutralization reaction is.
- Explain how neutralization reactions are used in acid-base titrations, and recall the vocabulary and setup of a titration from the class lab.
- Write a neutralization reaction for a given acid-base system and determine the concentration of one of the solutions given its volume and the volume and concentration of the other.
- Describe the dangers of hydrofluoric acid and the general treatment for accidental exposure
- Describe the causes and effects of acid rain and explain what has been done to reduce it.
- Labs: Acid Rainbow Demo; Invisible Fire; Titration of Vinegar; Acid Rain

General Skills from First Semester

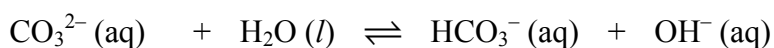
- Be able to use Table A to convert between any two units (i.e. J / kJ, L / mL)
- Mole conversions, density, naming and formulas of compounds

Chapter 18 Problem

- For the reaction, $2 \text{H}_2\text{S} (\text{g}) \rightleftharpoons 2 \text{H}_2 (\text{g}) + \text{S}_2 (\text{g})$,
 - Determine the equilibrium constant expression, K_{eq} , for this reaction.
 - Given $[\text{H}_2\text{S}] = 0.193 \text{ M}$, $[\text{H}_2] = 0.0035 \text{ M}$, and $[\text{S}_2] = 0.0035 \text{ M}$ at equilibrium, calculate the value for K_{eq} .

Chapter 19 Problems

- For following reaction, **circle the H^+ being transferred and draw an arrow** from the acid to the base *in the reactants* and **label** the acid (**A**), the base (**B**), the conjugate acid (**CA**), and the conjugate base (**CB**).



- A solution has $[\text{H}^+] = 1.0 \times 10^{-8} \text{ M}$. Determine the pH, $[\text{OH}^-]$, and pOH of this solution. Is this solution acidic, basic, or neutral?
- What is the pOH of a solution with $[\text{OH}^-] = 3.8 \times 10^{-10} \text{ M}$? What are the $[\text{H}^+]$ and pH of this solution? Is it acidic, basic, or neutral?
- Write the *balanced* neutralization equation for RbOH and malic acid, $\text{H}_2\text{C}_4\text{H}_4\text{O}_5$.
- In a titration, 30.0 mL of KOH solution is neutralized by 13.85 mL of 0.600 M HI solution.
 - Determine the number of moles of KOH neutralized in the reaction
 - Determine the concentration of the KOH solution

Answers to Chapter Problems
Chapter 18: #1. $K_{eq} = \frac{[\text{H}_2]^2 [\text{S}_2]}{[\text{H}_2\text{S}]^2} = 1.15 \times 10^{-6}$
Chapter 19: #2. A: H_2O ; B: OH^- ; CA: HCO_3^- ; CB: CO_3^{2-} ; #3. pH = 8, $[\text{OH}^-] = 1.0 \times 10^{-6} \text{ M}$, Basic; pOH = 6;
#4. pOH = 9.42, $[\text{H}^+] = 2.6 \times 10^{-5} \text{ M}$, pH = 4.58, Acidic; #5. $2 \text{Rb}^+ \text{C}_4\text{H}_4\text{O}_5^{2-} (\text{aq}) + 2 \text{H}_2\text{O} (\text{l}) \rightarrow 2 \text{Rb}^+ \text{OH}^- (\text{aq}) + \text{H}_2\text{C}_4\text{H}_4\text{O}_5 (\text{s})$; #6a. mol KOH = 0.00831 mol, #6b. $M_{\text{KOH}} = 0.277 \text{ M}$