

ACIDS
- a compound that produces hydrogen ions when dissolved
in water, the H is a +1 ion

NONOXYACID

A binary acid
(Has only 2 elements—no oxygen)

The first element is hydrogen and the
second element is a nonmetal

- The name of the acid begins with “**hydro**”
then the root of the nonmetal
and then ends with “**ic**”

Example:

HCl
(hydrogen chloride)
Hydrochloric acid

OXYACID

contains a polyatomic ion
(Has oxygen in the polyatomic ion)

The name of the acid begins with the root of the
polyatomic ion and then...

- if the ion ends with “**ate**”, the acid ends with “**ic**”
- if the ion ends with “**ite**” the acid ends with “**ous**”

(There is no “hydro” at start of name!)

Example:

H₂SO₄
(hydrogen sulfate)
sulfuric acid

Formula

Name

- | | | |
|-----|--|---------------------------|
| 1. | HBr | <u>hydrobromic acid</u> |
| 2. | H ₂ S | <u>hydrosulfuric acid</u> |
| 3. | H ₂ SO ₃ | <u>sulfurous acid</u> |
| 4. | HNO ₂ | <u>nitrous acid</u> |
| 5. | H ₃ PO ₄ | <u>phosphoric acid</u> |
| 6. | <u>HClO₃</u> | chloric acid |
| 7. | <u>HNO₃</u> | nitric acid |
| 8. | <u>HI</u> | hydroiodic acid |
| 9. | <u>HCH₃COO / CH₃COOH</u> | acetic acid |
| 10. | <u>HF</u> | hydrofluoric acid |