

Review Chapter 4: Study Guide Corrections

From the Exercise and Problems section, pp. 76-78

2. f. AgCl is a *weak* electrolyte because it is an insoluble salt.

$$22. \text{ b. } ? [\text{Mg}^{2+}] = 0.110 \text{ M MgCl}_2 \times \frac{1 \text{ mol Mg}^{2+}}{1 \text{ mol MgCl}_2} = \boxed{0.110 \text{ M Mg}^{2+}}$$

$$\text{c. } ? [\text{Cl}^-] = 0.110 \text{ M MgCl}_2 \times \frac{2 \text{ mol Cl}^-}{1 \text{ mol MgCl}_2} = \boxed{0.220 \text{ M Cl}^-}$$