

- Ch. 16.6: Solubility Equilibria & K_{sp}
- **Homework #16-6:** Problems pg. 723 #41, 45b, 46b, 50, 53, 54

45 (b) $[\text{Al}^{3+}] = 7.4 \times 10^{-8} \text{ M}$

46 (b) $K_{sp} = 1.8 \times 10^{-18}$

50 $s = 2.2 \times 10^{-4} \text{ mol/L}$

54 $[\text{F}^-] = 1.1 \times 10^{-4} \text{ M}$

$[\text{Sr}^{2+}] = 0.016 \text{ M}$

$[\text{NO}_3^-] = 0.076 \text{ M}$

$[\text{Na}^+] = 0.045 \text{ M}$