65   In water: $s = 1.4 \times 10^{-4}\ M$

In a buffer at pH $= 9.0$: $s = 0.12\ M$

67 $[OH^-] = 0.35\ M$

68 $Q = 2.0 \times 10^{-11}$, therefore, **a precipitate of Fe(OH)$_2$ will form**.

72 $[Cd^{2+}] = 1.1 \times 10^{-18}\ M$

$[CN^-] = 0.48\ M$

$[Cd(CN)_4^{2-}] = 4.2 \times 10^{-3}\ M$

74   **Overall: AgI(s) + 2NH$_3$(aq) $\rightleftharpoons$ Ag(NH$_3$)$_2^+$(aq) + I$^-$(aq) $K = K_{sp} \times K_f = 1.2 \times 10^{-9}$**

$s = 3.5 \times 10^{-5}\ M$