

From Zumdahl & Zumdahl 9<sup>th</sup> Edition p. 785:

99. The copper(I) ion forms a complex ion with  $\text{CN}^-$  according to the following equation:



- Calculate the solubility of  $\text{CuBr}(\text{s})$  ( $K_{\text{sp}} = 1.0 \times 10^{-5}$ ) in 1.0 L of 1.0 M NaCN
- Calculate the concentration of  $\text{Br}^-$  at equilibrium
- Calculate the concentration of  $\text{CN}^-$  at equilibrium