

WKS 1-4: Naming & Formulas of Ionic Compounds & Polyatomic Ions

<u>Name to Formula</u> Put charges on ions first, then make neutral. Omit charges (but keep subscripts) on final answer.	<u>Formula to Name</u> Name the cation and the anion.
1. calcium fluoride $\text{Ca}^{2+} \text{F}^{-} \rightarrow \text{CaF}_2$	2. MgBr_2 magnesium bromide
3. aluminum oxide $\text{Al}^{3+} \text{O}^{2-} \rightarrow \text{Al}_2\text{O}_3$	4. CsF cesium fluoride
5. magnesium nitride $\text{Mg}^{2+} \text{N}^{3-} \rightarrow \text{Mg}_3\text{N}_2$	6. Al_2S_3 aluminum sulfide
7. aluminum sulfite $\text{Al}^{3+} \text{SO}_3^{2-} \rightarrow \text{Al}_2(\text{SO}_3)_3$	8. LiCN lithium cyanide
9. magnesium hydroxide $\text{Mg}^{2+} \text{OH}^{-} \rightarrow \text{Mg}(\text{OH})_2$	10. $\text{Mg}(\text{HCO}_3)_2$ magnesium hydrogen carbonate
11. sodium hydrogen sulfate $\text{Na}^{+} \text{HSO}_4^{-} \rightarrow \text{NaHSO}_4$	12. CsNO_2 cesium nitrite
13. lithium iodate $\text{Li}^{+} \text{IO}_3^{-} \rightarrow \text{LiIO}_3$	14. $\text{Al}_2(\text{CO}_3)_3$ aluminum carbonate
15. calcium hypochlorite $\text{Ca}^{2+} \text{ClO}^{-} \rightarrow \text{Ca}(\text{ClO})_2$	16. NaBrO_2 sodium bromite
17. ammonium sulfate $\text{NH}_4^{+} \text{SO}_4^{2-} \rightarrow (\text{NH}_4)_2\text{SO}_4$	18. $(\text{NH}_4)_3\text{PO}_4$ ammonium phosphate
19. barium nitrate $\text{Ba}^{2+} \text{NO}_3^{-} \rightarrow \text{Ba}(\text{NO}_3)_2$	20. $\text{KC}_2\text{H}_3\text{O}_2$ potassium acetate