

**Balancing Equations Practice Worksheet 2**

*Balance the following equations:*

1. \_\_\_  $\text{HgO}(s) \xrightarrow{\Delta} \text{Hg}(l) + \text{O}_2(g)$
2. \_\_\_  $\text{Na}(s) + \text{H}_2\text{O} \rightarrow \text{NaOH}(aq) + \text{H}_2(g)$
3. \_\_\_  $\text{C}_{10}\text{H}_{16}(s) + \text{Cl}_2(g) \rightarrow \text{CCl}_4(g) + \text{HCl}(g)$
4. \_\_\_  $\text{C}_7\text{H}_6\text{O}_2(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}(g)$
5. \_\_\_  $\text{FeS}(s) + \text{O}_2(g) \rightarrow \text{Fe}_2\text{O}_3(s) + \text{SO}_2(g)$
6. \_\_\_  $\text{Fe}_2\text{O}_3(s) + \text{H}_2(g) \rightarrow \text{Fe}(s) + \text{H}_2\text{O}(g)$
7. \_\_\_  $\text{C}_2\text{H}_2(g) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}(g)$
8. \_\_\_  $\text{C}_7\text{H}_{16}(l) + \text{O}_2(g) \rightarrow \text{CO}_2(g) + \text{H}_2\text{O}(g)$
9. \_\_\_  $\text{KClO}_3(s) \rightarrow \text{KClO}_4(s) + \text{KCl}(s)$
10. \_\_\_  $\text{P}_4\text{O}_{10}(s) + \text{H}_2\text{O}(l) \rightarrow \text{H}_3\text{PO}_4(aq)$
11. \_\_\_  $\text{Fe}(s) + \text{H}_2\text{O}(l) \rightarrow \text{Fe}_3\text{O}_4(s) + \text{H}_2(g)$
12. \_\_\_  $\text{N}_2(g) + \text{O}_2(g) \rightarrow \text{N}_2\text{O}(g)$
13. \_\_\_  $\text{CO}_2(g) + \text{H}_2\text{O}(l) \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(s) + \text{O}_2(g)$
14. \_\_\_  $\text{SiCl}_4(g) + \text{H}_2\text{O}(l) \rightarrow \text{H}_4\text{SiO}_4(aq) + \text{HCl}(aq)$
15. \_\_\_  $\text{CO}_2(g) + \text{NH}_3(g) \rightarrow \text{CO}(\text{NH}_2)_2(s) + \text{H}_2\text{O}(l)$
16. \_\_\_  $\text{H}_2\text{SO}_4(aq) + \text{HI}(g) \rightarrow \text{H}_2\text{S}(g) + \text{I}_2(s) + \text{H}_2\text{O}(l)$
17. \_\_\_  $\text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + \text{Fe}(s)$
18. \_\_\_  $\text{Na}_2\text{CO}_3(aq) + \text{HCl}(aq) \rightarrow \text{NaCl}(aq) + \text{H}_2\text{O}(l) + \text{CO}_2(g)$
19. \_\_\_  $\text{Na}_2\text{O}_2(s) + \text{H}_2\text{O}(l) \rightarrow \text{NaOH}(aq) + \text{O}_2(g)$
20. \_\_\_  $\text{H}_3\text{AsO}_4(s) \xrightarrow{\Delta} \text{As}_2\text{O}_5(s) + \text{H}_2\text{O}(g)$
21. \_\_\_  $\text{FeCl}_3(aq) + \text{NH}_4\text{OH}(aq) \rightarrow \text{Fe}(\text{OH})_3(s) + \text{NH}_4\text{Cl}(aq)$
22. \_\_\_  $\text{Ca}_3(\text{PO}_4)_2(s) + \text{SiO}_2(s) \rightarrow \text{P}_4\text{O}_{10}(s) + \text{CaSiO}_3(s)$
23. \_\_\_  $\text{N}_2\text{O}_5(g) + \text{H}_2\text{O}(l) \rightarrow \text{HNO}_3(aq)$
24. \_\_\_  $\text{NaOH}(aq) + \text{Cl}_2(g) \rightarrow \text{NaCl}(aq) + \text{NaClO}(aq) + \text{H}_2\text{O}(l)$
25. \_\_\_  $\text{V}_2\text{O}_5(s) + \text{HCl}(aq) \rightarrow \text{VOCl}_3(aq) + \text{H}_2\text{O}(l)$