

Ionic Compounds WKS II
Formula & Nomenclature Practice

NAME _____
Period _____ Date _____

<u>Name to Formula</u> (Put charges on ions first!!! Then make neutral.)	<u>Formula to Name</u> (Don't forget roman numerals for metals with more than one possible charge.)
1. tin(IV) chromate	11. PbO ₂
2. zinc carbonate	12. FeSO ₄
3. mercury (II) chloride	13. Ni(NO ₂) ₃
4. cobalt(II) sulfide	14. PdCO ₃
5. iron (III) oxide	15. CuCl ₂
6. vanadium(V) hydroxide	16. Mo ₂ S ₃
7. manganese (IV) carbonate	17. NbI ₅
8. copper (II) nitrate	18. RuBr ₃
9. chromium(VI) oxide	19. Zn(OH) ₂
10. iridium(IV) phosphate	20. AgNO ₃

21. What does the Roman numeral in the name of a cation indicate? How can it be determined from the formula?

22. When is it necessary to use a Roman numeral in the name of a cation? What cations do NOT need a Roman numeral?