

**Ionic Compounds WKS I**  
**Formula & Nomenclature Practice**

NAME Answer Key  
 Period \_\_\_\_\_ Date \_\_\_\_\_

<u>Name to Formula</u> Put charges on ions first, then make neutral. Omit charges (but keep subscripts) on final answer.	<u>Formula to Name</u> Name the cation and the anion.
1. calcium fluoride $\text{Ca}^{2+} \text{F}^{-} \rightarrow \text{CaF}_2$	16. $\text{MgBr}_2$ <b>magnesium bromide</b>
2. aluminum oxide $\text{Al}^{3+} \text{O}^{2-} \rightarrow \text{Al}_2\text{O}_3$	17. $\text{CsF}$ <b>cesium fluoride</b>
3. magnesium nitride $\text{Mg}^{2+} \text{N}^{3-} \rightarrow \text{Mg}_3\text{N}_2$	18. $\text{Al}_2\text{S}_3$ <b>aluminum sulfide</b>
4. sodium phosphide $\text{Na}^{+} \text{P}^{3-} \rightarrow \text{Na}_3\text{P}$	19. $\text{Rb}_3\text{P}$ <b>rubidium phosphide</b>
5. lithium selenide $\text{Li}^{+} \text{Se}^{2-} \rightarrow \text{Li}_2\text{Se}$	20. $\text{Na}_3\text{N}$ <b>sodium nitride</b>
6. cesium oxide $\text{Cs}^{+} \text{O}^{2-} \rightarrow \text{Cs}_2\text{O}$	21. $\text{LiI}$ <b>lithium iodide</b>
7. strontium oxide $\text{Sr}^{2+} \text{O}^{2-} \rightarrow \text{SrO}$	22. $\text{NH}_4\text{Cl}$ <b>ammonium chloride</b>
8. aluminum sulfite $\text{Al}^{3+} \text{SO}_3^{2-} \rightarrow \text{Al}_2(\text{SO}_3)_3$	23. $\text{LiCN}$ <b>lithium cyanide</b>
9. magnesium hydroxide $\text{Mg}^{2+} \text{OH}^{-} \rightarrow \text{Mg}(\text{OH})_2$	24. $\text{Mg}(\text{HCO}_3)_2$ <b>magnesium hydrogen carbonate</b>
10. sodium hydrogen sulfate $\text{Na}^{+} \text{HSO}_4^{-} \rightarrow \text{NaHSO}_4$	25. $\text{CsNO}_2$ <b>cesium nitrite</b>
11. lithium iodate $\text{Li}^{+} \text{IO}_3^{-} \rightarrow \text{LiIO}_3$	26. $\text{Al}_2(\text{CO}_3)_3$ <b>aluminum carbonate</b>
12. calcium hypochlorite $\text{Ca}^{2+} \text{ClO}^{-} \rightarrow \text{Ca}(\text{ClO})_2$	27. $\text{K}_2\text{Cr}_2\text{O}_7$ <b>potassium dichromate</b>
13. strontium carbonate $\text{Sr}^{2+} \text{CO}_3^{2-} \rightarrow \text{SrCO}_3$	28. $\text{NaBrO}_2$ <b>sodium bromite</b>
14. ammonium sulfate $\text{NH}_4^{+} \text{SO}_4^{2-} \rightarrow (\text{NH}_4)_2\text{SO}_4$	29. $(\text{NH}_4)_3\text{PO}_4$ <b>ammonium phosphate</b>
15. barium nitrate $\text{Ba}^{2+} \text{NO}_3^{-} \rightarrow \text{Ba}(\text{NO}_3)_2$	30. $\text{KC}_2\text{H}_3\text{O}_2$ <b>potassium acetate</b>