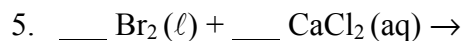
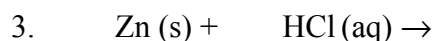
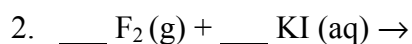
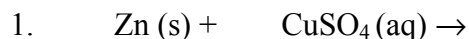


WKS – Honors Chem
Single Replacement Reaction Products

NAME _____
Period _____ **Date** _____

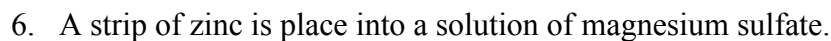
Part A

Determine the products and write the balanced equation for each for each of these reactions assuming that the reaction proceeds. Then use the Standard Reduction Potentials in Chart H to determine if the reaction does proceed. If the reaction cannot occur, write “NR” after it. Finally write the complete and net ionic equations. Remember—elements have no charge. Aqueous strong acids dissociate into H^+ and their anion in the complete ionic equation.



Part B

Determine the reactant formulas for the following scenarios, then complete the reactions as above.



7. An iron nail is placed into a solution of nickel(II) nitrate. (Fe forms the Fe^{2+} ion)

8. A solid chunk of calcium is placed into water.

9. Chlorine gas is bubbled through a solution of strontium bromide.

10. Cadmium pellets are placed into a solution of perchloric acid.