

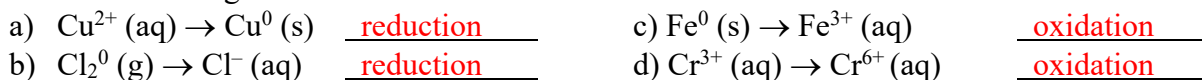
**WKS – Chem Honors
Oxidation and Reduction (Redox)**

Name Answer Key
Period _____ Date _____

1) When a substance is oxidized, its charge (**increases, decreases**). Thus, when a substance is oxidized it (**gains, loses**) electrons.

2) When a substance is reduced, its charge (**increases, decreases**). Thus, when a substance is reduced it (**gains, loses**) electrons.

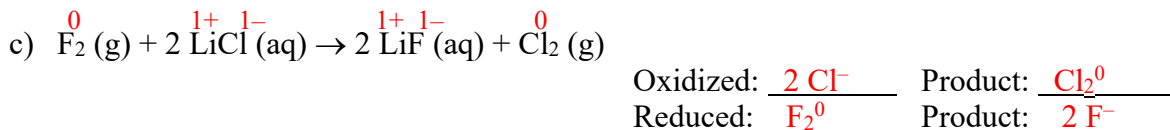
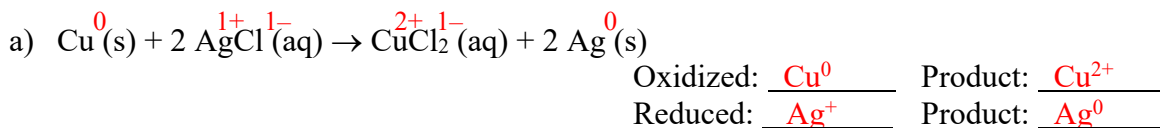
3) Label these changes as either oxidation or reduction:



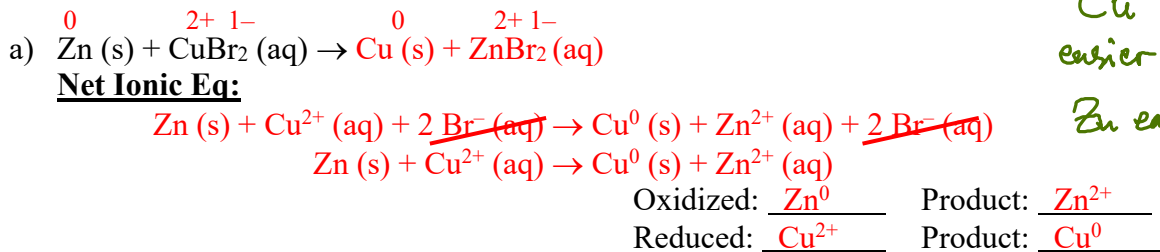
4) Put in all charges (oxidation numbers) in the following substances. (Remember: The charge on pure elements is zero. Also, when H is in a compound, its charge is +1)



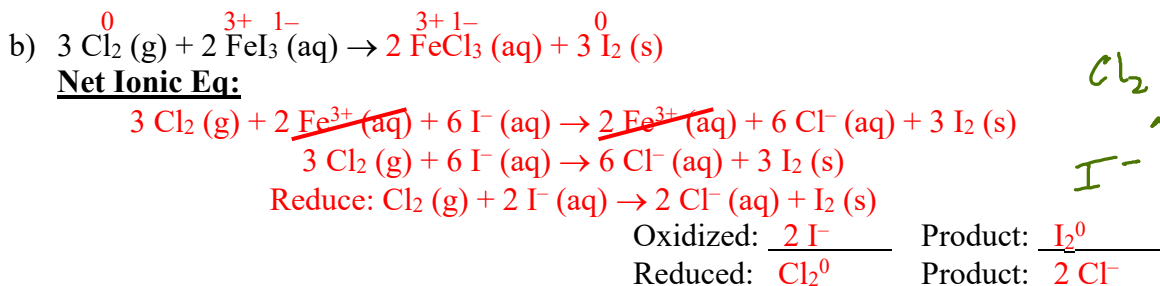
5) Put in all charges (oxidation numbers). Then indicate which substance is being oxidized, which is being reduced, and what their products are. [Remember LEO-GER!]



6) Put in all charges (oxidation numbers) and then write the net ionic equation. (Only write substances whose charges change, not spectator ions.) Then indicate which substance is being oxidized, which is being reduced, and what their products are. [Remember LEO-GER!]



$\text{Cu}^{2+} \rightarrow \text{Cu}$
easier to reduce than Zn^{2+}
 Zn easier to oxidize than Cu



Cl_2 better oxidizing agent than I_2
 I^- better reducing agent than Cl^-