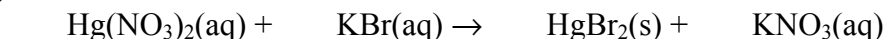


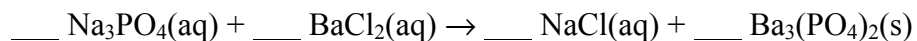
WKS – Honors
Solution Stoichiometry

Name _____
Period _____

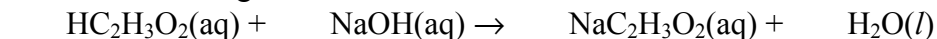
- 1) How many mL of 0.348 M $\text{Hg}(\text{NO}_3)_2$ are needed to fully react with 15.0 mL of 0.485 M KBr according to the following unbalanced reaction?



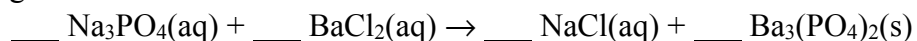
- 2) What volume of 0.325 M Na_3PO_4 would be needed to fully react with 25.0 mL of 0.480 M BaCl_2 by the following unbalanced reaction?



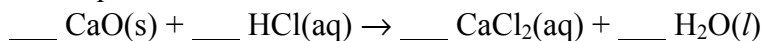
- 3) Calculate the volume of 0.2250 M $\text{HC}_2\text{H}_3\text{O}_2$ (acetic acid) solution needed to neutralize 25.19 mL of 0.4295 M NaOH in the following unbalanced reaction.



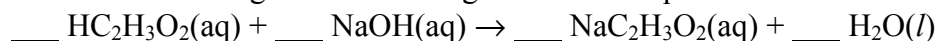
- 4) What volume of 0.325 M Na_3PO_4 would be needed to precipitate 25.00 g $\text{Ba}_3(\text{PO}_4)_2$ with excess BaCl_2 by the following unbalanced reaction?



- 5) How many grams of CaO are required for complete reaction with the HCl in 275 mL of a 0.523 M HCl solution? The unbalanced equation for the reaction is:



- 6) 34.57 mL of HC₂H₃O₂ solution of unknown concentration is used to neutralize 25.19 mL of NaOH with concentration 0.4295 M according to the following unbalanced equation:



- a) How many moles of acetic acid are used?

- b) What is the concentration of the acetic acid solution, in M?

- 7) When 321 mL of HCl solution of unknown concentration reacts with Na₂CO₃, it forms NaCl, water, and 11.1 g of CO₂ according to the following unbalanced equation:



- a) How many moles of HCl are used in the reaction?

- b) What is the concentration of the HCl solution, in M?