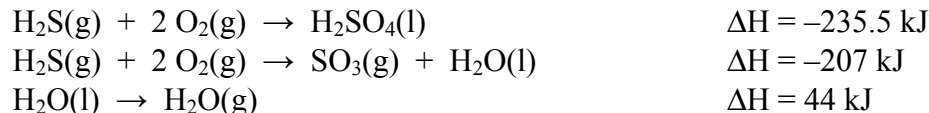


WKS-Honors  
Hess's Law

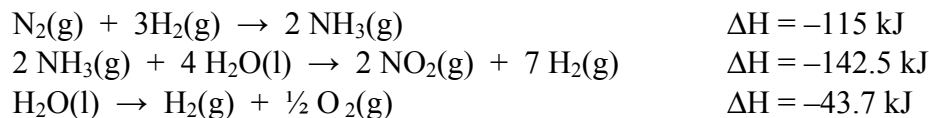
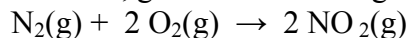
Name \_\_\_\_\_  
Period \_\_\_\_\_ Date \_\_\_\_\_

Use Hess's Law to find the enthalpy changes for the stated reactions given the component reactions and their enthalpy changes. Use the space given to rewrite the modified component reactions and show that they add up to the desired reaction.

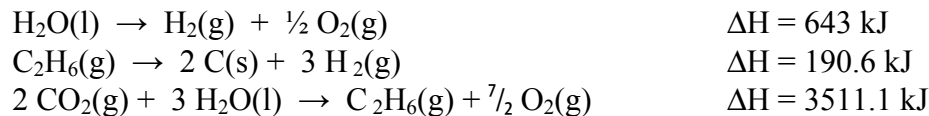
- 1) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



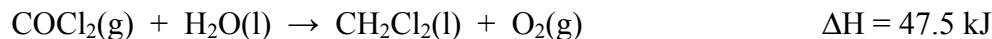
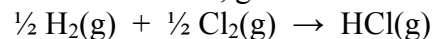
- 2) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



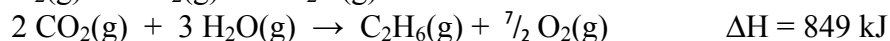
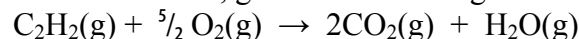
- 3) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



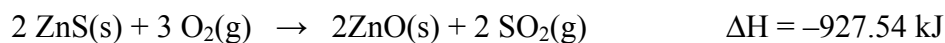
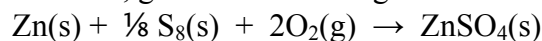
4) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



5) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



6) Find the  $\Delta H$  for the reaction below, given the following reactions and subsequent  $\Delta H$  values:



Answers: 1) 73 kJ; 2) -83 kJ; 3) 886 kJ; 4) -230 kJ; 5) -705 kJ; 6) -976.03 kJ